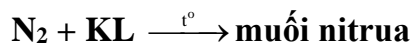
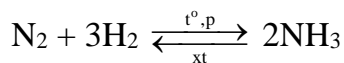


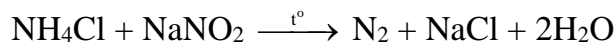
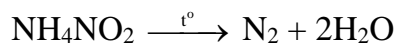
# MỘT SỐ PHƯƠNG TRÌNH ÔN TẬP KTTT HỌC KÌ I

## NITƠ

Tính oxi hóa:

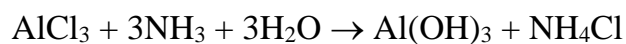
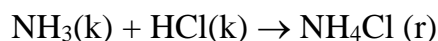
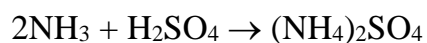


Điều chế:

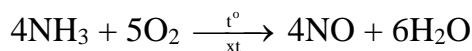
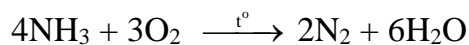


## AMONIAC

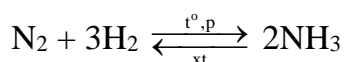
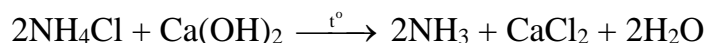
Tính bazơ yếu:



Tính khử:

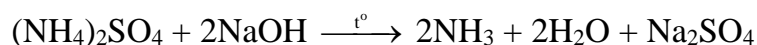


Điều chế:

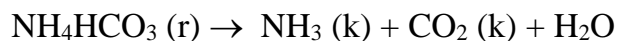
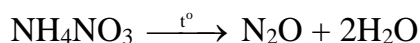
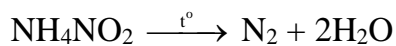
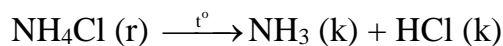


## MUỐI AMONI

Tác dụng với dung dịch kiềm:



Phản ứng nhiệt phân:

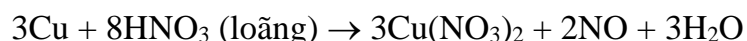


## AXIT NITRIC

1. Tính axit:

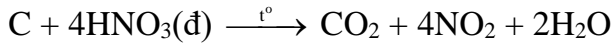
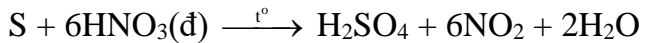
2. Tính oxi hóa mạnh:

☞ Tác dụng Kim loại:

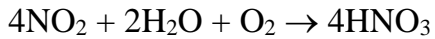
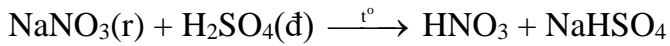


●\* **Fe, Al không tác dụng với HNO<sub>3</sub> đặc nguội**

☞ **Tác dụng Phi kim:**

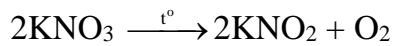


Điều chế:

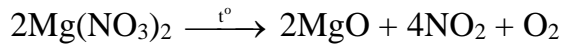


## MUỐI NITRAT

❖ Đối với muối nitrat của kim loại từ K → trước Mg tạo ra **muối nitrit + O<sub>2</sub>**



❖ Đối với muối nitrat của kim loại từ Mg → Cu tạo ra **oxit KL + NO<sub>2</sub> + O<sub>2</sub>**



❖ Đối với muối nitrat của kim loại từ Ag → về sau tạo ra **kim loại + NO<sub>2</sub> + O<sub>2</sub>**

